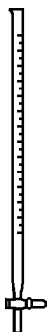


- 1) A. 0.0563 g
Chemistry, Ch 17 #2
- 2) B. 0.3010 mg
Chemistry, Ch 17 #4
- 3) D. 0.029 gram
Chemistry, Ch 17 #6
- 4) D. 7.44 g
Chemistry, Ch 17 #8
- 5) B. 60.87 g
Chemistry, Ch 17 #10
- 6) C. 5,797.9 g
Chemistry, Ch 17 #12
- 7) A. 0.08 °C/g
Chemistry, Ch 17 #14
- 8) C. 2.58 cm²
Chemistry, Ch 17 #16
- 9) A. 1.65
Chemistry, Ch 17 #22
- 10) C. 14.6
Chemistry, Ch 17 #24
- 11) A. 16.9%
Chemistry, Ch 17 #26
- 12) C. 12.5%
Chemistry, Ch 17 #28
- 13) B. 37.8% water
Chemistry, Ch 17 #30
- 14) B. 
Chemistry, Ch 17 #36

15)

D.



Chemistry, Ch 17 #38

16)

D.



Chemistry, Ch 17 #40

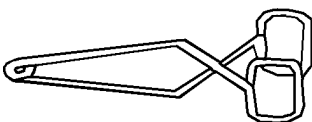
17)

C. crucible tongs

Chemistry, Ch 17 #42

18)

B.



Chemistry, Ch 17 #44

19)

A. Erlenmeyer flask

Chemistry, Ch 17 #46

20)

D. watch glass

Chemistry, Ch 17 #48

21)

A. Insufficient oxygen is mixing with the gas.

Chemistry, Ch 17 #50

22)

B. titration

Chemistry, Ch 17 #52

23)

C. 1.25 cm

Chemistry, Ch 17 #54

24)

D. 545.20 g

Chemistry, Ch 17 #56

25)

D. 75.7 mL

Chemistry, Ch 17 #58

26)

D. 31.28 mL

Chemistry, Ch 17 #60

- 27) A. 2.7 g/cm^3
Chemistry, Ch 17 #62
- 28) C. 2.7
Chemistry, Ch 17 #64
- 29) A. 3.4 g/mL
Chemistry, Ch 17 #66
- 30) B. $60. \text{ cm}^3$
Chemistry, Ch 17 #68
- 31) B. 22.4 mL
Chemistry, Ch 17 #70
- 32) C. The temperature of the solution decreases.
Chemistry, Ch 17 #74
- 33) D. salt and water
Chemistry, Ch 17 #76
- 34) D. The dissolving of salt *A* was exothermic and the dissolving of salt *B* was endothermic.
Chemistry, Ch 17 #78
- 35) A. The gas is hydrogen.
Chemistry, Ch 17 #80
- 36) A. Na_2CO_3
Chemistry, Ch 17 #82
- 37) C. Place the open test tube in a beaker of water, then heat the water in the beaker.
Chemistry, Ch 17 #84
- 38) D. 4
Chemistry, Ch 17 #90
- 39) C. 53°C
Chemistry, Ch 17 #92
- 40) B. 53°C
Chemistry, Ch 17 #94
- 41) B. *B*
Chemistry, Ch 17 #96

42) (a) 14.6 g/cm^3

WORK SHOWN: $D = \frac{M}{V}$, $D = \frac{37.89 \text{ g}}{2.60 \text{ cm}^3} = 14.6 \text{ g/cm}^3$;

(b) SAMPLE ANSWER: The actual value for the density of gold is 19.320 g/cm^3 ; therefore, she should conclude that the sample is not pure gold.

Chemistry, Ch 17 #104

43) (a) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$;

(b) 1.0 M

WORK SHOWN: $M_A V_A = M_B V_B$, $(2.0 \text{ M})(10.02 \text{ mL}) = M_B(20.00 \text{ mL})$, $M_B = 1.0 \text{ M}$;

(c) SAMPLE ANSWER: Averaging three trials gives a more accurate result. OR When a single trial is run, you have no way of determining if an error was made or if an experimental result was inaccurate.

Chemistry, Ch 17 #105